

## DEMOVERSION OF TESTING

For tasks 1-3, use the following row of chemical elements.

1)Be      2)H      3)F      4)Li      5)Si

1. Determine from the elements indicated above which atoms have one electron missing for completion of electronic shell.

Write down the numbers of the selected elements in the answer field

Answer:

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2. Select three elements that are in the Periodic Table are in the same period. Arrange the selected chemical elements in increasing order of their atomic radius. Write down the numbers of the selected elements in the required sequence in the answer field.

Answer:

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3. Select two compounds from indicated above between which a hydrogen bond occurs .

- 1) methane
- 2) silane
- 3) ethanol
- 4) phosphine
- 5) formic acid

Write down the numbers of the selected compounds in the answer field

Answer:

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4. Establish a correspondence between the formula of the substance and the class / group to which this substance belongs.

FORMULA OF SUBSTANCE

CLASS or GROUP

- A)  $\text{HMnO}_4$
- B)  $\text{KHSiO}_3$
- C)  $\text{CO}$

- 1) acidic oxide
- 2) non-salt-forming oxide
- 3) acidic salt
- 4) acid

Write down the selected numbers in the table under the corresponding letters

Answer:

A	B	C

5. Two test tubes with a solution of chromium (III) chloride are given. A solution of a weak electrolyte X was added to one of them, and a solution of a strong electrolyte Y was added to the other. As a result, the formation of a precipitate was observed in each of the test tubes. From the proposed list, select substances X and Y that can enter into the described reactions.

- |                     |                    |
|---------------------|--------------------|
| 1) hydrogen bromide | 4) calcium bromide |
| 2) ammonia          | 5) silver nitrate  |
| 3) hydrogen iodide  |                    |

Write down the numbers of the selected substances in the answer field

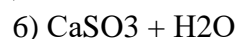
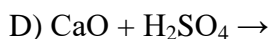
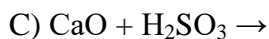
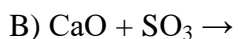
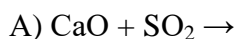
Answer:

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6. Establish a correspondence between the reagents and the product(s) of their interaction.

REAGENTS

PRODUCT(S)



Write down the selected numbers in the table under the corresponding letters

Answer:

A	B	C	D

7. The following scheme for the transformation of substances is given



Determine which of these substances are substances X and Y.

- 1)  $\text{H}_2$
- 2)  $\text{H}_2\text{O}_2$
- 3)  $\text{H}_2\text{O}$
- 4)  $\text{ZnSO}_4$
- 5)  $\text{K}_2\text{SO}_4$

Write down the numbers of the selected substances under the corresponding letters in the table.

Answer:

X	Y

8. Establish a correspondence between the name of the substance and the class / group to which this substance belongs.

NAME OF SUBSTANCE

CLASS or GROUP

- A) methyl benzoate
- B) ethylene glycol
- C) alanine

- 1) alcohol
- 2) amino acid
- 3) ester
- 4) nitro compound

Write down the selected numbers in the table under the corresponding letters

Answer:

A	B	C

9. Select two substances that are homologues

- 1) benzene
- 2) styrene
- 3) toluene
- 4) glycerol
- 5) glycine

Write down the numbers of the selected elements in the answer field

Answer:

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10. Choose two substances that decolorize bromine water

- 1) propane
- 2) ethylene
- 3) divinyl
- 4) cyclopentane
- 5) isobutane

Write down the numbers of the selected elements in the answer field

Answer:

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11. From the proposed list, select two substances with which ethanol reacts

- 1) potassium
- 2) aluminum hydroxide
- 3) nitrogen

- 4) potassium bicarbonate  
5) formic acid

Write down the numbers of the selected elements in the answer field

Answer:

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12. The following scheme for the transformation of substances is given



Determine which of these substances are substances X and Y.

- 1) H<sub>2</sub> (cat)
- 2) H<sub>2</sub>O<sub>2</sub>
- 3) H<sub>2</sub>O(Hg<sup>2+</sup>)
- 4) KMnO<sub>4</sub>(H<sup>+</sup>)
- 5) O<sub>2</sub> (cat)

Write down the selected numbers in the table under the corresponding letters

Answer:

X	Y

13. Establish a correspondence between the formula of salt and the products of electrolysis of its aqueous solution, which were released on inert electrodes

SUBSTANCE FORMULA	ELECTROLYSIS PRODUCTS
A) Al(NO <sub>3</sub> ) <sub>3</sub>	1) metal and oxygen
B) Hg(NO <sub>3</sub> ) <sub>2</sub>	2) hydrogen and oxygen
C) KCl	3) hydrogen and halogen
D) CsNO <sub>3</sub>	4) metal and halogen
	5) metal and nitric oxide (IV)

Write down the selected numbers in the table under the corresponding letters

Answer:

A	B	C	D

14. For listed substances, determine the nature of the environment of their aqueous solutions having the same concentration (mol / l).

- 1)  $K_2SO_4$
- 2)  $ZnCl_2$
- 3)  $Na_2S$
- 4)  $HClO_4$

Write down the numbers of substances in ascending order of the pH value of their aqueous solutions.

Answer: \_\_\_\_\_

15. Establish a correspondence between the external influence on the system and a shift in chemical equilibrium in the system.

REACTION EQUATION	CHANGE OF EXTERNAL CONDITIONS
A) $N_{2(g)} + 3H_{2(g)} \leftrightarrow 2NH_{3(g)} + Q$	1) increase in temperature and hydrogen concentration
B) $H_2O_{(g)} + CH_4 \leftrightarrow 3H_{2(g)} + CO_{(g)} - Q$	2) decrease in temperature and hydrogen concentration
B) $H_{2(g)} + I_{2(TB)} \leftrightarrow 2HI_{(g)} - Q$	3) increase in temperature and decrease in hydrogen concentration
Г) $C_2H_6_{(g)} \leftrightarrow C_2H_4_{(g)} + H_2_{(g)} - Q$	4) decrease in temperature and increase in hydrogen concentration

Write down the selected numbers in the table under the corresponding letters

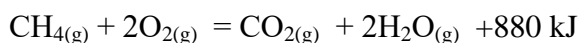
Answer:

A	B	C	D

16. 5 ml of water and 10 g of sodium nitrate were added to 200 g of a solution with a mass fraction of sodium nitrate of 12%. What is the mass fraction of salt in the resulting solution?

Answer: \_\_\_\_\_ % (Write down the number to the nearest tenth.)

17. In accordance with the thermochemical equation



during the combustion of 56 l of methane (normal conditions), heat is released in the amount \_\_\_\_\_ kJ

Answer: \_\_\_\_\_ kJ (Write down the integer number)

18. Select substances (potassium hypochlorite, potassium hydroxide, iron (III) sulfate, chromium (III) oxide, magnesium oxide, sodium iodide) between which a redox reaction is possible in an alkaline environment

Write the molecular equation for the reaction with the correct coefficients

Create an electronic balance:

Specify the oxidizing agent \_\_\_\_\_

Specify a reducing agent \_\_\_\_\_

19. For electrolysis (on inert electrodes), 390 g of a 15% sodium chloride solution were taken. After the mass of the solution decreased by 21.9 g, the process was stopped. To the resulting solution was added 160 g of a 20% solution of copper (II) sulfate.

Write reaction equations with correctly placed coefficients

\_\_\_\_\_

Calculate the number of moles of starting substances (reactants) taking part in the reaction

\_\_\_\_\_

Calculate the mass of the final solution

\_\_\_\_\_

Calculate the mass fraction of silver nitrate in the resulting solution

\_\_\_\_\_

20. When burning 2.55 g of organic matter, 3.36 liters of carbon dioxide (normal conditions) and 3.15 g of water were obtained. It is known that this compound does not enter into an esterification reaction and does not interact with metallic sodium. It can be obtained in one step from isopropanol

Calculate the number of moles of combustion products

\_\_\_\_\_

Calculate the number of moles of carbon, hydrogen, etc. atoms. Find the molar ratio of atoms

\_\_\_\_\_

Write the molecular formula of an organic substance

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Write the structural formula of an organic substance

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Write the reaction equation for obtaining this substance from isopropanol

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