

Reactivity series of metals

Li	K	Ba	Ca	Na	Mg	Al	Be	Mn	Cr	Zn	Fe	Cd	Co	Ni	Sn	Pb	H ₂	Sb	Cu	Hg	Ag	Pt	Au
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Table of substance's solubility at 20⁰ C

SOL- soluble substances ;INSL- insoluble substances; SLSL – slightly soluble substances

	K ⁺	Na ⁺	Ba ²⁺	Ca ²⁺	Mg ²⁺	Al ³⁺	Cr ³⁺	Fe ²⁺	Mn ²⁺	Zn ²⁺	Ag ⁺	Hg ²⁺	Cu ²⁺	Pb ²⁺	Fe ³⁺
OH⁻	sol	sol	sol	sIsl	sIsl	insl	insl	insl	insl	insl	-	-	insl	insl	insl
Cl⁻	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	insl	sol	sol	M	sol
S²⁻	sol	sol	sol	sIsl	sol	-	-	insl	insl	insl	insl	insl	insl	insl	-
SO₃²⁻	sol	sol	insl	insl	insl	-	-	insl	insl	insl	insl	insl	insl	insl	-
SO₄²⁻	sol	sol	insl	sIsl	sol	sol	sol	sol	sol	sol	sIsl	sol	sol	insl	sol
PO₄³⁻	sol	sol	insl	insl	insl	insl	insl	insl	insl	insl	insl	insl	insl	insl	insl
CO₃²⁻	sol	sol	insl	insl	insl	-	-	insl	insl	insl	insl	insl	insl	insl	insl
SiO₃²⁻	sol	sol	insl	insl	insl	insl	insl	insl	insl	insl	insl	-	sIsl	sIsl	insl
NO₃⁻	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol
CH₃COO⁻	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol	sol

Measure of solubility in the circumstances is the substance content in a saturated solution.

Solubility varies to a wide extent:

$\omega = 10\%$ (very soluble), $\omega = 1\%$ (slightly soluble),

$\omega = 0,01\%$ (practically insoluble).

The solubility of a substance depends on temperature: the solid substance's solubility increases together with the increase in temperature.